

Physical Science

Elements Quiz #1

1. Supply names or symbols as needed:

- a. K - potassium
- b. Mercury - Hg
- c. Na - sodium
- d. Gold - Au
- e. W - tungsten
- f. Silver - Ag
- g. Fe - iron
- h. Tin - Sn
- i. Sb - antimony
- j. Copper - Cu
- k. Pb - lead

Part of atom	Charge	mass	Location
1. proton	2. +	3. 1 Amu	4. nucleus
5. neutron	6. 0 (no chg)	7. 1 Amu	8. nucleus
9. electron	10. -	11. $\frac{1}{1837}$ Amu	12. levels orbitals - shells

3. What is an isotope? Give an Example.

Atoms of the same element that have different numbers of neutrons. (ex: carbon 12, carbon 13, carbon 14)

5. On the Periodic Table each element is represented by 2 numbers, the whole number is the atomic number. The decimal number is the atomic mass. To calculate the number of ~~protons~~<sup>neutrons</sup> you must subtract the atomic number from the atomic mass.

6. Which number tells the number of protons?

The atomic number (whole number) indicates the number of protons of the element.

7. Supply missing information and write the element name to the left of each element.

Atomic number	Atomic Mass	# of Protons	# of electrons	# of neutrons
6	12	6	6	6
8	16	8	8	8
17	35	17	17	18